

**SIMPLE IONS**

Complete the following table. Note that the name of a nonmetallic ion ends with *-ide*, while the name for a metallic ion uses the full name of the metal.

Ion Name	Ion Symbol	Number of Protons	Number of Valence Electrons	Number of Electrons Lost or Gained	Same Number of Electrons as an Ion as What Noble Gas?
e.g. fluoride	F <sup>-</sup>	9		gained one	neon
1.		53			
2.		16		gained two	
3. potassium				lost one	
4.	Ca <sup>2+</sup>				
5.		35			
6.	Sr <sup>2+</sup>				
7.	H <sup>+</sup>				(none)
8.		8		gained two	
9.		12		lost two	
10. aluminum			3		
11.		34			
12.	H <sup>-</sup>				
13. lithium				lost one	
14.	Rb <sup>+</sup>				
15.		17	7		

Note that the symbols for ions are written, for example, as Ca<sup>2+</sup>. This <sup>2+</sup> symbol is recommended by IUPAC (The international Union of Pure and Applied Chemistry). The incorrect designation <sup>+2</sup> is sometimes confused as indicating a positive number or an oxidation number, rather than indicating a positive charge.